Errata second printing:

pg 38:   should read: $n_1 = n_o x$, $n_2 = n_o x^2$, $n_3 = n_o x^3$,   etc.

pg 110:   problem 5.22 should read "…using the bond dissociation energies of H—H, C—H, C—C, and C=C bonds given in Table 5.1."

pg. 152   section 8.11

$$S = \Delta_r S^o - R \ln P_{H_2O(g)} / bar \quad should\ be \quad \Delta_r S = \Delta_r S^o - R \ln P_{H_2O(g)} / bar$$

pg. 156   For all problems in this chapter, assume T = 25 $^\circ$C if the temperature is not given.

pg. 157   problem 8.25(a) should read $P_4(g) + 5\,O_2(g) \longrightarrow P_4O_{10}(s)$

pg. 230   problem 12.11 should read "(a) Using thermodynamic data for both Hg(ℓ) and Hg(g) at 25 $^\circ$C (which can be found at webbook.nist.gov) determine the vapor pressure of mercury at 25 $^\circ$C. Also use this data to estimate the standard boiling point of mercury."

pg 289   "Answers to selected problems" 7.23: the entropy change in going from state B to C should be -6.87 x 10^{-23} J/K.